

Electron Configuration Test

Name _____

Period _____ Date _____

Write the complete electron configuration for the following elements.

1) Fe

2) Ca

3) Ag

Write the noble gas configuration for the following elements.

4) Br _____ 5) Po _____ 6) Mo _____

7) a) Write the orbital diagram for selenium (Se).

7b) How many electrons are in the highest energy level of selenium (Se)? _____

8) How many electrons can one orbital hold? _____

9) What is the maximum number of electrons that each of the following sublevels can hold?

s _____ p _____ d _____ f _____

10) How many unpaired electrons are in the last sublevel of phosphorous? _____

11) How many electrons are in the outermost **energy level** of arsenic (As) _____

12) Using the electron configurations, identify the following elements:

[Ar]4s²3d⁶ _____ [Xe]6s²4f¹⁴5d¹⁰6p⁵ _____

13) Electrons that are closer to the nucleus have less energy than those further from the nucleus. **True** **False**